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created during ordinary chemical reactions  
or physical changes. law of definite  
proportions.

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Modern Chemistry Chapter 3. law of conservation of mass. law of definite proportions. law of multiple proportions. atom. States that matter can be neither destroyed nor created during.... The fact that a chemical compounds contains the same elements.... if two or more different compounds are composed of the same ...

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Chapter 3. Structure of the Atom. Mass and Isotopes. Development of Atomic Theory. The Mole and Mole Concept. Chapter 4. Matter as Waves. Quantum Model. Electron Configuration.

Chapter 3 - Chemistry

Modern Chemistry Chapter 3. Atoms: The Building Blocks of Matter. law of conservation of mass- mass is neither created nor destroyed during ordinary

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chemical reactions or physical changes. e.g.  
 $20 \text{ g A} + 20 \text{ g B} \rightarrow 40 \text{ g AB}$ .

## Modern Chemistry Chapter 3 Atoms: The Building Blocks of ...

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CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. In cathode-ray tubes, the cathode ray is emitted from the negative electrode, which is called the cathode . 2. The smallest unit of an element that can exist either alone or in

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molecules containing the

## 3 Atoms: The Building Blocks of Matter

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Blocks of Matter law of conservation of  
mass mass is neither created nor destroyed  
during ordinary chemical reactions or  
physical changes e.g.  $20\text{ g A} + 20\text{ g B} \rightarrow 40\text{ g AB}$   
law of definite proportions- a chemical  
compound contains the same elements in  
exactly the same proportions by mass  
regardless of the size of the sample or the  
source of the compound e.g. If 10 grams of  
A combine with 20 grams of B to form  
compound AB, how ...

## Modern Chemistry Chapter 3 - Modern Chemistry Chapter 3...

CHAPTER 3 TEST Class Atoms: The  
Building Blocks of Matter MULTIPLE  
CHOICE On the line at the left of each  
statement, write the letter of the choice best

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completes the statement or answers the question. The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays were composed of a. energy

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