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Answers Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and

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frequency are inversely related; in other words, as one quantity increases, the other decreases.

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Key Concepts: Terms in this
set (57) Dalton. The atom is
a tiny, indestructible
particle with no internal

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Answers. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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Speed of light 5.2 Bohr's
Model of the Atom/Quantum
Mechanical Model of the Atom
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Models of the Atom
Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the

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planets move around the sun.
Rutherford's model fails to
explain why objects change
color when heated.

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Section 5.2 - Electron
Arrangement in Atoms The
electron configuration of an
atom is the arrangement of
the electrons. There are 3
rules that govern the
electron configuration:
Aufbau's principle, Pauli

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Answers Exclusion principle, and Hund's rule.

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Answers * * * * * The
electromagnetic spectrum
consists of radiation over a
broad band of wavelengths.
The visible light portion is
very small. It is in the
 10^{-7}m wavelength range and
 $10^{15}\text{ Hz (s}^{-1}\text{)}$ frequency

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Answers Mechanical Model. model of the atom we believe today that involves the probability of finding an electron in a certain position. What is the maximum number of f orbitals in any single energy level

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Answers
in an atom ? 7. Electrons in
the same orbital.

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unique characteristic of
that atom. You will compare

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Answers and contrast the Bohr and
quantum mechanical

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Configurations for Elements
in Period Three Table 5-4

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Answers Figure 5-19. This sublevel diagram shows the order in which the orbitals are usually filled. The proper sequence for the first seven orbitals is 1s, 2s, 2p, 3s, 3p, 4s, and 3d. Chapter 5
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Rutherford used existing
ideas about the atom and

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Answers proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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Rutherford's model fails to explain why objects change color when heated.

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p orbital hold? Chapter 5:
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2. Light and Quantized

Energy (5.1)

- The study of light led to the development of the quantum mechanical model.

- Light is a kind of electromagnetic

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radiation EM).

All move at 3.00×10^8 m/s (c) Speed of light.

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