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s  $4.0 \text{ m} \sin 30^\circ$  g d ( ) (2) ( ) ( ) For a correct numerical answer . 1 point .  $u = 5.3 \text{ m/s}$ . Alternate solution (kinematics method) Alternate points . For determination of the linear acceleration in terms of mass and frictional force : 1 point .  $a_{FM} = 52 \text{ f}$  from work shown in part (b)

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acceleration in terms of mass and frictional force : 1 point .  
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(9 points)  $8 \text{H}^+(\text{aq}) + 4 \text{Cl}^-(\text{aq}) + \text{MnO}_4^-(\text{aq}) \rightarrow 2 \text{Cl}_2(\text{g}) + \text{Mn}^{2+}(\text{aq}) + 4 \text{H}_2\text{O}(\text{l})$   
 $\text{Cl}_2(\text{g})$  can be generated in the laboratory by reacting potassium permanganate with an acidified solution of sodium chloride. The net-ionic equation for the reaction is given above. (a) A 25.00 mL sample of 0.250 M NaCl reacts completely with excess KMnO

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AP CHEMISTRY 2010 SCORING GUIDELINES

Score: 10. The response earned 1 point in part (a) for

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correctly plotting the means on a bar graph. The response earned 1 point in part (a) for appropriate labels, units, and scaling. The response earned 1 point in part (a) for correctly plotting the error bars.

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### AP Biology 2017 FRQ 1 Student Samples

Briefly explain your answer. | An unknown solid has a melting point of 112-114°C. Based on data from Table 1 above and the mixed melting point technique of this experiment, you can narrow down the possibilities to three compounds. List the compounds b) A mixed melting point of the unknown with acetanilide gives a melting point of 113-114°C.

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1. To observe the reaction to be studied, transfer 10-mL of 1.5% (0.44 M) H<sub>2</sub>O<sub>2</sub> into a 50-mL glass beaker and add 1 mL of the freshly made catalase solution. The bubbles coming from the reaction mixture are O<sub>2</sub>, which results from the breakdown of H<sub>2</sub>O<sub>2</sub> by catalase. Be sure to keep the freshly made catalase solution on ice at all times.

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Experiment 24: Electrochemistry: Voltaic Cells - AP Chem ...  
AP CHEMISTRY BY SATELLITE 9 - 7 INSTRUCTOR LAB  
MANUAL EXPERIMENT 9: SOLUBILITIES Pre-lab

Questions: Answer these questions before coming to class.  
They introduce you to several important ideas that you will  
use in this experiment. You must turn-in this exercise before  
you will be allowed to begin the experiment.

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